

## Oxygen Trap System (OTS®) for the Reduction of Residual Oxygen in NETZSCH Thermal Analyzers

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### Introduction

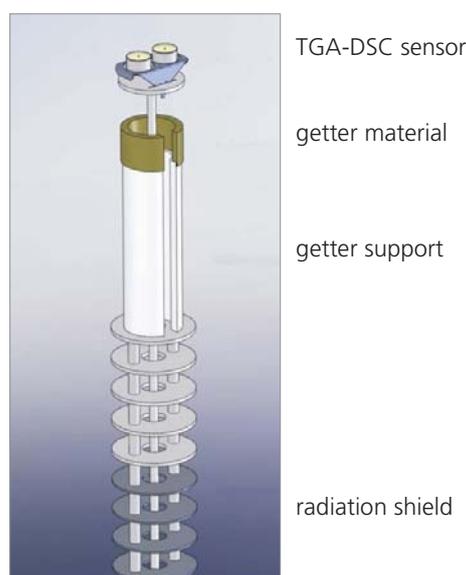
The presence of residual oxygen is a well known problem in thermal analysis (please refer to the terminology given in DIN 51 005). When samples are to be investigated under inert gas conditions using for example, nitrogen, argon or helium as a purge gas, the presence of residual oxygen is decisive in most cases because possible oxidation of the sample would lead to undesired results and incorrect interpretations.

Metallic samples oxidizing on the surface exhibit an exothermic DSC signal as well as an increase in sample mass. The oxidation can also be responsible for a shift in the phase transformation temperatures. Polymers or composites containing organics would partially combust in the presence of residual oxygen which would falsify the measurement result during the nominally pyrolytic decomposition.

Residual oxygen in thermal analyzers is typically minimized by evacuating, backfilling and purging of the apparatus with a highly pure inert gas. This process should be repeated several times in order to minimize the oxygen concentration. The most important prerequisite for the lowest possible oxygen concentration is of course a vacuum-tight instrument. Thus, the concentration of residual oxygen depends on the vacuum-tightness of the thermal analyzer, of the gas lines and of the gas connections as well as on the purity of the inert purge gas. Additional cleaning of the purge gas outside of the instrument can be helpful but does usually not yield completely satisfactory results.

### The OTS® System

The OTS® system allows for an additional, efficient *in-situ* reduction of the oxygen concentration at the sample site. Figure 1 shows the OTS® system installed in a simultaneous thermal analyzer (STA = TGA + DSC): Below the sample and

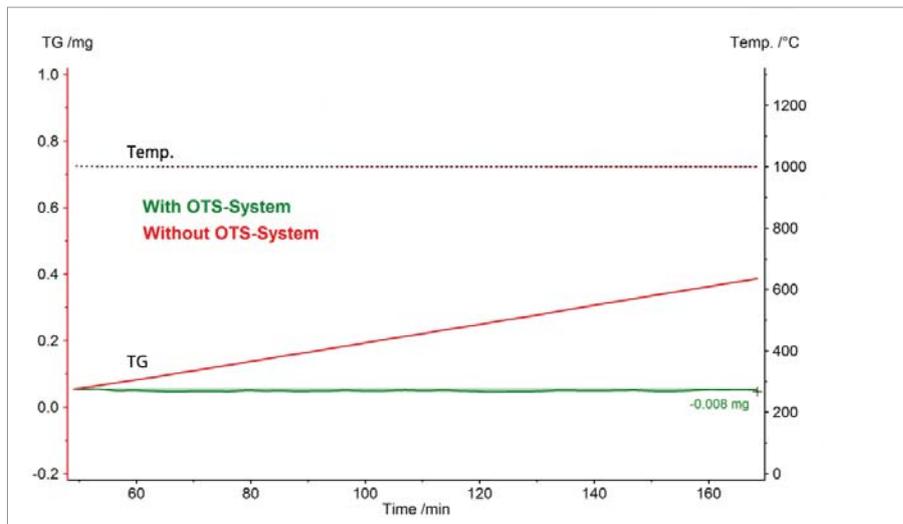


1 OTS® system for the reduction of residual oxygen installed in a simultaneous thermal analyzer (STA)

reference crucible and thus in the hot zone of the instrument is a highly temperature-resistant getter material which absorbs residual oxygen at sufficiently high temperatures. The getter material is positioned by a ceramic getter support, which is also highly temperature-resistant and does not react with the getter material. Both parts, the getter material and the ceramic getter support, are placed on the radiation shield of the TGA-DSC sample carrier.

The rotational symmetry ensures that the OTS® system is not in direct contact with the sample carrier. And due to the slitted design of the getter material and the ceramic getter support, the OTS® system can be easily mounted or removed. The inert purge gas flowing upwards first comes into contact with the getter material and then with the sample. Therefore, residual oxygen which is present in the purge gas is completely absorbed by the getter material and can thus not reach the sample.

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2 Mass change (TGA) obtained for two zirconium samples at 1000°C. One sample (green) was measured with the OTS® system, the other without (red).

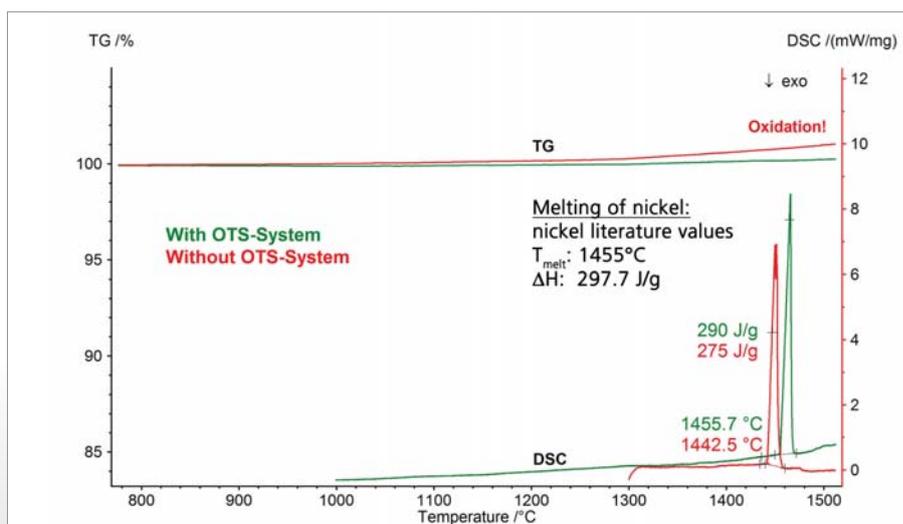
### Results and Discussion

Two TGA measurements on zirconium are compared in figure 2; one with the OTS® system and the other without. Both measurements were carried out in a dynamic helium atmosphere with a nominal purity of 99.996%. The samples were held isothermally at 1000°C for about 2 hours. Without the OTS® system, the sample mass increased at a constant rate, finally arriving at 0.33 mg. This mass gain reflecting the oxidation of the sample was avoidable with the OTS® system: The sample mass remained almost constant. From these results, it can be estimated that the OTS® system

reduces the residual oxygen concentration at the sample site to below ~1 ppm.

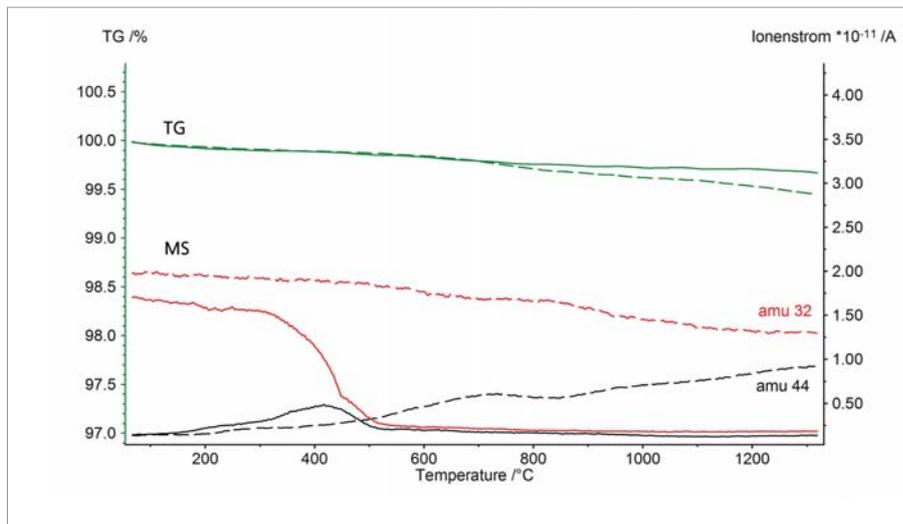
Another example demonstrating the benefit of the OTS® system is shown in figure 3. Two nickel samples were investigated by means of a simultaneous thermal analyzer. Argon purge gas with a purity of 99.996% was used in both cases. The literature melting point of nickel of 1455°C is often employed for thermometry at high temperatures. Nickel, however, is very sensitive to oxidation which

can lead to an undefined reduction of the melting point and thus to incorrect thermometry. This can be seen in the measurement without the OTS® system: The sample oxidized, resulting in an increase of the TGA curve due to the mass gain. The DSC melting peak occurred already at 1443°C which is 12°C lower than the literature value. The melting enthalpy of 275 J/g is also significantly lower than the literature value of approx 300 J/g. Correct results, corresponding to literature values, were obtained with the OTS® system: The DSC melting peak was detected at 1455°C and the enthalpy of melting was 290 J/g.



3 TGA-DSC results obtained for two nickel samples. One sample (green) was measured with the OTS® system, the other without (red).

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4 Mass change (TGA) and mass spectrometer signals (ion current for mass numbers 32 and 44) obtained for two graphite samples. One sample was measured with the OTS® system (solid lines), the other without (dashed lines).

Due to the OTS® system, the sample did not oxidize significantly. This can be seen via the horizontal TGA curve, which means that the sample mass was constant during the course of the experiment.

Finally, depicted in figure 4 are two TGA-MS measurements on two graphites which were again carried out in argon purge gas with a purity of 99.996%. The slight mass loss below ~600°C is most probably due to volatile hydrocarbons, whereas the mass loss at higher temperatures observed without the OTS® system reflects the partial burn-up of the graphite due to residual oxygen (dashed lines): The mass spectrometer detected an increase of the signal with mass number 44 which is due to CO<sub>2</sub> evolution; the gradual decrease of the signal with mass number 32 reflects the corresponding consumption of residual oxygen.

With the OTS® system, the sample mass remained practically constant above ~600°C which means that the sample showed no further oxidation (solid lines). There was also no CO<sub>2</sub> evolution detected in that temperature range. From the oxygen signal (mass number 32), it can also be concluded that the OTS® system starts to absorb residual oxygen above ~300°C and reduces the oxygen concentration to a minimum above ~500°C.

### Conclusion

The OTS® oxygen trap systems can be employed with various thermal analyzers (DSC, TGA, STA, DIL). It removes the traces of residual oxygen in the gas atmosphere inside the instrument to concentrations well below 1 ppm.